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2 → CaSiO<sub>3</sub> B. 3Fe<sub>2</sub>O<sub>3</sub> + CO → 2Fe<sub>3</sub>O<sub>4</sub> + CO<sub>2</sub> C. FeO + SiO<sub>2</sub> → FeSiO<sub>3</sub> D. FeO → Fe + 1/2 O<sub>2</sub> A. A B. D C. C And D D. A And D Answer: C Solution: In Metallurgy Of Iron, CaO Is Used As Flux Which Is Used To Remove The Impurities Of SiO<sub>2</sub>, CaO +

$\text{SiO}_2 \rightarrow \text{CaSiO}_3$  . Also Here  $\text{Fe}_2\text{O}_3$  Is Reduced By CO To  $\text{Fe}_3\text{O}_4$  Which Is Further Reduced To ... Apr 3th, 2024

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