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Environmental Chemistry 2.1 INTRODUCTION 2.2 STOICHIOMETRY ...42 Chapter 2 Environmental Chemistry Section 2.2 Stoichiometry 43 We Already Know That There Are 44 Grams Per Mole Of CO" So We Do Not Need To Recalculate That. Two Moles Of Butane ($2 \text{ Mol} \times 58 \text{ G/mol} = 116 \text{ G}$) Yields 8 Moles Of Carbon Dioxide ($8 \text{ Mol} \times 44 \text{ G/mol} = 352 \text{ G CO}_2$) , So We Can Set Up The Following Proportion: Jul 3th, 2024.

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Chemical Equations Another Variation On Conversion Factors 1. Add The Use Of Mole Ratios As Conversion Factors 2. Instead Of Converting Single Compounds/elements From Moles Mass # Particles, Use Information About One Compound/element In A Chemical Reaction To Determine The Jun 1th, 2024Stoichiometry: Mole-Mole Problems - Mr. V's Chemistry SiteChemistry IF8766 Page 62 Instructional Fair, Inc. Title: Microsoft Word - Pg 62 - Stoichiome Feb 3th, 2024.

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Actual yield = 1.45g Theoretical yield = 5.95g $\text{KClO}_4 \times 1 \text{ mol } 122.55 \text{ g} \times 2 \times 2 \times 1 \text{ KClO}_4 \times 32.00 \text{ g } 1 \text{ mol} = 3.11 \text{ g O}_2$ % Yield = $\frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$

$4 \text{ FeCO}_3 + \text{O}_2 \rightarrow 2 \text{ Fe}_2\text{O}_3 + 4 \text{ CO}_2$ Percent purity = 62.8%? g pure FeCO_3

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