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$N\text{O} + n\text{e}^- \rightleftharpoons \text{NRR}$ (2.1.36) The Free Energy Of The Cell Reaction Is Given By, $\Delta G = \Delta G_0 + RT \ln \frac{a_{\text{R}}}{a_{\text{N}}}$ $\text{H}^+ + \text{H} \rightleftharpoons \text{H}^+$ (2.1.38) Which Since $\Delta G = -nFE_{\text{rxn}}$ And $\frac{a_{\text{H}^+}}{a_{\text{H}^+}} = 1$ Becomes, $E = E_0 + \frac{RT}{nF} \ln \frac{a_{\text{R}}}{a_{\text{O}}}$. (2.1.40) This Expression Is Called The Nernst Equation. This Relation Can Jan 1th, 2024

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