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A) [Co(H₂O)₆]²⁺ Co +2 E) [FeO]₄²⁻ Fe +6 I) K₂[CoCl₄] Co +2 B) [CrCl₆]³⁻ Cr +3 F) [Mn(CN)₆]⁴⁻ Mn +2 J) K₃[AuF₆] Au +3 C) [Co(NH₃)₆]Cl₂ Co +2 G) [Ni(CO)₄] Ni 0 K) (NH₄)₂[IrCl₆] Ir +4 D) [Co(NH₃)₅Cl]Cl₂ Co +3 H) [Ni(EDTA)]²⁻ Ni +2 L) Na[Mn(CO)₅]

Mn -1 TASK 2 – Writing Half Equations A) $2I^- \rightarrow I_2 + 2e^-$
– B ... Mar 3th, 2024

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It Burns ... Apr 2th, 2024

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Metal Carbonate \rightarrow Metal Oxide + Carbon Dioxide (on Heating)
 $CaCO_3 \rightarrow CaO + CO_2$ TASK 4 – WRITING
BALANCED EQUATIONS 1) Balance The Following
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Chromatography . Chromatography Is Usually Introduced As A Technique For Separating And/or Identifying The ... All Chromatographic Methods Require One Static Part (the Stationary Phase) And One Moving Part (the Mobile Phase). The Techniques Rely On One Of The Following Phenomena: Adsorption; ... The Result Is A Separation Into Bands Containing ...

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Examples Of Chromatography Chromatography In Real Life

- Choose One Of The Testing Markers And Draw A Thick Line Near The Bottom Of The Filter Paper - About $\frac{1}{4}$ Inch From The Bottom.
- Pour A Small Amount Of Water Into The Large Cup And Then Hang The Paper Strip In The

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An Experiment To Find The Activation Energy For The
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Iodine Monochloride And Chlorine In An Equilibrium
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